



## INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

(51) International Patent Classification <sup>6</sup> : <b>H01M 8/10, C08J 5/22, H01M 8/02, 2/16</b>		A1	(11) International Publication Number: <b>WO 96/29752</b>
			(43) International Publication Date: 26 September 1996 (26.09.96)
(21) International Application Number: <b>PCT/US96/03804</b>		(US). RAJENDRAN, Govindarajulu [IN/US]; Apartment No. 11, English Village Apartments, Newark, DE 19711-5925 (US).	
(22) International Filing Date: 20 March 1996 (20.03.96)		(74) Agent: HENDRICKSON, John, S.; E.I. du Pont de Nemours and Company, Legal/Patent Records Center, 1007 Market Street, Wilmington, DE 19898 (US).	
(30) Priority Data:		(81) Designated States: CA, CN, JP, SG, US, European patent (AT, BE, CH, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE).	
08/407,146 20 March 1995 (20.03.95) US			
08/412,791 29 March 1995 (29.03.95) US			
60/003,736 14 September 1995 (14.09.95) US			
(60) Parent Applications or Grants		Published With international search report. Before the expiration of the time limit for amending the claims and to be republished in the event of the receipt of amendments.	
(63) Related by Continuation			
US 08/407,146 (CIP)			
Filed on 20 March 1995 (20.03.95)			
US 08/412,791 (CIP)			
Filed on 29 March 1995 (29.03.95)			
US 60/003,736 (CIP)			
Filed on 14 September 1995 (14.09.95)			
(71) Applicant (for all designated States except US): E.I. DU PONT DE NEMOURS AND COMPANY [US/US]; 1007 Market Street, Wilmington, DE 19898 (US).			
(72) Inventors; and			
(75) Inventors/Applicants (for US only): GROT, Walther, Gustav [DE/US]; 4 Byron Court, Chadds Ford, PA 19317-9204			
(54) Title: MEMBRANES CONTAINING INORGANIC FILLERS AND MEMBRANE AND ELECTRODE ASSEMBLIES AND ELECTROCHEMICAL CELLS EMPLOYING SAME			
(57) Abstract			
<p>Cation exchange membranes (26), made from polymer having cation exchange groups and containing inorganic filler, exhibit reduced fuel crossover for fuel cells (10) employing direct feed organic fuels (12) such as methanol and, when inorganic proton conductor is employed, enhanced proton conductivity.</p>			

**FOR THE PURPOSES OF INFORMATION ONLY**

Codes used to identify States party to the PCT on the front pages of pamphlets publishing international applications under the PCT.

AM	Armenia	GB	United Kingdom	MW	Malawi
AT	Austria	GE	Georgia	MX	Mexico
AU	Australia	GN	Guinea	NE	Niger
BB	Barbados	GR	Greece	NL	Netherlands
BE	Belgium	HU	Hungary	NO	Norway
BF	Burkina Faso	IE	Ireland	NZ	New Zealand
BG	Bulgaria	IT	Italy	PL	Poland
BJ	Benin	JP	Japan	PT	Portugal
BR	Brazil	KE	Kenya	RO	Romania
BY	Belarus	KG	Kyrgyzstan	RU	Russian Federation
CA	Canada	KP	Democratic People's Republic of Korea	SD	Sudan
CF	Central African Republic	KR	Republic of Korea	SE	Sweden
CG	Congo	KZ	Kazakhstan	SG	Singapore
CH	Switzerland	LJ	Liechtenstein	SI	Slovenia
CI	Côte d'Ivoire	LK	Sri Lanka	SK	Slovakia
CM	Cameroon	LR	Liberia	SN	Senegal
CN	China	LT	Lithuania	SZ	Swaziland
CS	Czechoslovakia	LU	Luxembourg	TD	Chad
CZ	Czech Republic	LV	Latvia	TG	Togo
DE	Germany	MC	Monaco	TJ	Tajikistan
DK	Denmark	MD	Republic of Moldova	TT	Trinidad and Tobago
EE	Estonia	MG	Madagascar	UA	Ukraine
ES	Spain	ML	Mali	UG	Uganda
FI	Finland	MN	Mongolia	US	United States of America
FR	France	MR	Mauritania	UZ	Uzbekistan
GA	Gabon			VN	Viet Nam

TITLE

MEMBRANES CONTAINING INORGANIC FILLERS AND  
MEMBRANE AND ELECTRODE ASSEMBLIES AND  
ELECTROCHEMICAL CELLS EMPLOYING SAME

5 FIELD OF THE INVENTION

This invention relates to ion exchange  
membranes and membrane and electrode assemblies useful  
in electrochemical cells and fuel cells such as fuel  
cells employing direct feed organic fuels such as  
10 methanol.

BACKGROUND OF THE INVENTION

A variety of known electrochemical cells fall  
within a category of cells often referred to as solid  
polymer electrolyte (SPE) cells. An SPE cell typically  
15 employs a membrane of an ion exchange polymer which  
serves as a physical separator between the anode and  
cathode while also serving as an electrolyte. SPE  
cells can be operated as electrolytic cells for the  
production of electrochemical products or they may be  
20 operated as fuel cells for the production of electrical  
energy. The most well known fuel cells are those which  
operate with gaseous fuels such as hydrogen and with a  
gaseous oxidant, usually pure oxygen or oxygen from  
air, and those fuel cells using direct feed organic  
25 fuels such as methanol.

In some SPE cells including many fuel cells, a  
cation exchange membrane is employed and protons are  
transported across the membrane as the cell is  
operated. Such cells are often referred to as proton  
30 exchange membrane (PEM) cells. For example, in a cell  
employing the hydrogen/oxygen couple, hydrogen  
molecules (fuel) at the anode are oxidized donating  
electrons to the anode, while at the cathode the oxygen  
(oxidant) is reduced accepting electrons from the  
35 cathode. The  $H^+$  ions (protons) formed at the anode  
migrate through the membrane to the cathode and combine  
with oxygen to form water. In many fuel cells, the  
anode and/or cathode are provided by forming a layer of

electrically conductive, catalytically active particles, usually also including a polymeric binder, on the proton exchange membrane and the resulting structure (sometimes also including current collectors) is referred to as a membrane and electrode assembly or MEA.

Membranes made from a cation exchange polymer such as perfluorinated sulfonic acid polymer have been found to be particularly useful for MEA's and electrochemical cells due to good conductivity and good chemical and thermal resistance which provides long service life before replacement. However, increased proton conductivity is desired for some applications, particularly for fuel cells which operate at high current densities.

In fuels cells which employ direct feed organic fuels such as methanol, a problem with known cells has been the so-called crossover of fuel through the membrane. The term "crossover" refers to the undesirable transport of fuel through the membrane from the fuel electrode or anode side to the oxygen electrode or cathode side of the fuel cell. After having been transported across the membrane, the fuel will either evaporate into the circulating oxygen stream or react with the oxygen at the oxygen electrode.

The fuel crossover diminishes cell performance for two primary reasons. Firstly, the transported fuel cannot react electrochemically and, therefore, contributes directly to a loss of fuel efficiency (effectively a fuel leak). Secondly, the transported fuel interacts with the cathode i.e., the oxygen electrode, and lowers its operating potential and hence the overall cell voltage. The reduction of cell voltage lowers specific cell power output and also reduces the overall efficiency. Therefore, it is especially desirable to provide a cation exchange

membrane for use in a fuel cell which has a low fuel crossover rate.

#### SUMMARY OF THE INVENTION

The invention provides improved cation exchange  
5 membranes and also membrane and electrode assemblies  
and electrochemical cells which employ the membranes.  
A cation exchange membrane in accordance with the  
invention is made from polymer having cation exchange  
groups and having inorganic filler dispersed therein.  
10 In one preferred form of the invention, the polymer in  
the bulk of the membrane is substantially free of metal  
catalysts. Preferably, the inorganic filler is an  
inorganic proton conductor, more preferably an  
inorganic proton conductor selected from the group  
15 consisting of particle hydrates and framework hydrates.  
Most preferably, the inorganic proton conductor is  
selected from the group consisting of titanium dioxide,  
tin and hydrogen mordenite, oxides and phosphates of  
zirconium, and mixtures thereof. It is also preferred  
20 for the inorganic proton conductor to have a  
conductivity of at least about  $10^{-4}$  S/cm. It is also  
preferred for the inorganic proton conductor to  
comprise about 2% to about 30%, most preferably, about  
5% to about 25% by weight of the membrane.  
25 Preferably, the cation exchange groups of the  
polymer are selected from the group consisting of  
sulfonate, carboxylate, phosphonate, imide, sulfonimide  
and sulfonamide. In one preferred form of the  
invention, highly fluorinated polymer with sulfonate  
30 cation exchange groups is employed. It has been found  
to be particularly advantageous for the polymer to  
comprise a highly fluorinated carbon backbone with its  
side chains represented by the formula  
 $-(OCF_2CFR_f)_a-OCF_2CFR'_fSO_3X$ , wherein  $R_f$  and  $R'_f$  are  
35 independently selected from F, Cl or a perfluorinated  
alkyl group having 1 to 10 carbon atoms,  $a = 0, 1$  or  $2$ ,  
and  $X$  is H, an alkali metal, or  $NH_4$ .

The invention also provides a fuel cell for operation with a direct feed organic fuel which includes an anode compartment containing an anode, a cathode compartment containing a cathode and a membrane serving as a separator and electrolyte between the anode and cathode compartments. The membrane comprises a polymer having cation exchange groups and which has inorganic filler dispersed therein. A preferred organic fuel is methanol.

Fuel cells employing direct feed organic fuels such as methanol made with the membranes of the present invention have reduced crossover of methanol fuel, providing enhanced electrical performance. Preferred membranes in accordance with the invention having inorganic proton conductor dispersed therein have enhanced proton conductivity and electrochemical cells such as fuel cells made with membranes of this invention have enhanced electrical performance.

Membranes of this invention also have enhanced mechanical properties due to incorporation of inorganic fillers, e.g., increased stiffness.

The invention also provides a membrane and electrode assembly comprising a cation exchange membrane of polymer having cation exchange groups with an electrode formed on at least one of its surfaces. The electrode comprises a layer of electrically conductive, catalytically active particles and binder polymer having cation exchange groups. In accordance with the invention, the binder polymer has inorganic filler dispersed therein.

The invention also provides a process for making cation exchange membranes containing dispersed inorganic filler. A process in accordance with the invention comprises forming a membrane of polymer having cation exchange groups and dispersing inorganic filler in the membrane by *in situ* precipitation. Preferred forms of the process in accordance with the invention are useful for dispersing the inorganic

proton conductors titanium dioxide and zirconium hydrogen phosphate in the membrane.

#### BRIEF DESCRIPTION OF THE DRAWING

Figure 1 is a schematic drawing which depicts the structure of a membrane and electrode assembly (MEA) in accordance with the present invention.

#### DETAILED DESCRIPTION OF THE INVENTION

##### Membrane Polymers

A membrane in accordance with the invention is made of a polymer having cation exchange groups which can transport protons across the membrane. The cation exchange groups are preferably selected from the group consisting of sulfonate, carboxylate, phosphonate, imide, sulfonimide and sulfonamide groups. Various known cation exchange polymers can be used including polymers and copolymers of trifluoroethylene, tetrafluoroethylene, styrene-divinyl benzene,  $\alpha,\beta,\beta$ -trifluorostyrene, etc., in which cation exchange groups have been introduced.  $\alpha,\beta,\beta$ -trifluorostyrene polymers useful for the practice of the invention are disclosed in U.S. Patent 5,422,411.

In a preferred form of the invention, the polymer comprises a polymer backbone and recurring side chains attached to the backbone with the side chains carrying the cation exchange groups. For example, copolymers of a first fluorinated vinyl monomer and a second fluorinated vinyl monomer having a side cation exchange group or a cation exchange group precursor can be used, e.g., sulfonyl fluoride groups ( $-\text{SO}_2\text{F}$ ) which can be subsequently hydrolyzed to sulfonic acid groups. Possible first monomers include tetrafluoroethylene, hexafluoropropylene, vinyl fluoride, vinylidene fluoride, trifluoroethylene, chlorotrifluoroethylene, perfluoro(alkyl vinyl ether), and mixtures thereof. Possible second monomers include a variety of fluorinated vinyl ethers with cation exchange groups or precursor groups.

Preferably, the polymer in accordance with the invention has a polymer backbone which is highly fluorinated and the ion exchange groups are sulfonate groups. The term "sulfonate groups" is intended to refer either to sulfonic acid groups or alkali metal or ammonium salts of sulfonic acid groups. "Highly fluorinated" means that at least 90% of the total number of halogen and hydrogen atoms are fluorine atoms. Most preferably, the polymer backbone is perfluorinated. It is also preferable for the side chains to be highly fluorinated and, most preferably, the side chains are perfluorinated.

A class of preferred polymers for use in the present invention include a highly fluorinated, most preferably perfluorinated, carbon backbone and the side chain is represented by the formula  $-(OCF_2CFR_f)_a-OCF_2CFR'_fSO_3X$ , wherein  $R_f$  and  $R'_f$  are independently selected from F, Cl or a perfluorinated alkyl group having 1 to 10 carbon atoms,  $a = 0, 1$  or  $2$ , and  $X$  is H, an alkali metal, or  $NH_4$ . The preferred polymers include, for example, polymers disclosed in U.S. Patents 4,358,545 and 4,940,525. Most preferably, polymer comprises a perfluorocarbon backbone and said side chain is represented by the formula  $-O-CF_2CF(CF_3)-O-CF_2CF_2SO_3X$ , wherein  $X$  is H, an alkali metal, or  $NH_4$ . Polymers of this type are disclosed in U.S. Patent 3,282,875

The equivalent weight of the cation exchange polymer can be varied as desired for the particular application. For the purposes of this application, equivalent weight is defined to be the weight of the polymer in sulfonic acid form required to neutralize one equivalent of NaOH. In the case where the polymer comprises a perfluorocarbon backbone and the side chain is the salt of  $-O-CF_2-CF(CF_3)-O-CF_2-CF_2-SO_3X$ , the equivalent weight preferably is 800-1500, most preferably 900-1200. The equivalent weight of the



polymers disclosed in U.S. Patents 4,358,545 and 4,940,525 is preferably somewhat lower, e.g., 600-1300.

5 In the manufacture of membranes using polymer which has a highly fluorinated polymer backbone and sulfonate ion exchange groups, membranes are typically formed from the polymer in its sulfonyl fluoride form since it is thermoplastic in this form and conventional techniques for making films from thermoplastic polymer can be used. Alternately, the polymer may be in 10 another thermoplastic form such as by having  $-SO_2X$  groups where X is  $CH_3$ ,  $CO_2$ , or a quaternary amine. Solution film casting techniques using suitable solvents for the particular polymer can also be used if desired.

15 A film of the polymer in sulfonyl fluoride form can be converted to the sulfonate form (sometimes referred to as ionic form) by hydrolysis using methods known in the art. For example, the membrane may be hydrolyzed to convert it to the sodium sulfonate form 20 by immersing it in 25% by weight NaOH for about 16 hours at a temperature of about  $90^\circ C$  followed by rinsing the film twice in deionized  $90^\circ C$  water using about 30 to about 60 minutes per rinse. Another possible method employs an aqueous solution of 6-20% of 25 an alkali metal hydroxide and 5-40% polar organic solvent such as dimethyl sulfoxide with a contact time of at least 5 minutes at  $50-100^\circ C$  followed by rinsing for 10 minutes. After hydrolyzing, the membrane can be converted if desired to another ionic form by 30 contacting the membrane in a bath containing a 1% salt solution containing the desired cation or, to the acid form, by contacting with an acid and rinsing. For fuel cell use, the membrane is usually in the sulfonic acid form.

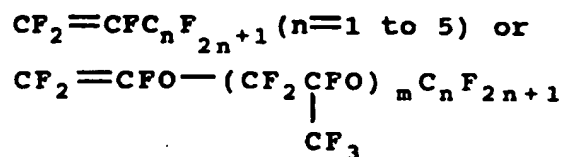
35 If desired, the membrane may be a laminated membrane of two polymers such as two highly fluorinated polymers having different ion exchange groups and/or different ion exchange capacities. Such membranes can

be made by laminating two films or co-extruding a film with the two polymer layers. In addition, the membrane may be made of a blend of two or more polymers such as two or more highly fluorinated polymers having  
 5 different ion exchange groups and/or different ion exchange capacities.

The thickness of the membrane can be varied as desired for a particular electrochemical cell application. Typically, the thickness of the membrane  
 10 is generally less than about 250  $\mu\text{m}$ , preferably in the range of about 25  $\mu\text{m}$  to about 150  $\mu\text{m}$ .

The membrane may optionally include a porous support for the purposes of improving mechanical properties, for decreasing cost and/or other reasons.  
 15 The porous support of the membrane may be made from a wide range of components. The porous support of the present invention may be made from a hydrocarbon such as a polyolefin, e.g., polyethylene, polypropylene, polybutylene, copolymers of those materials, and the  
 20 like. Perhalogenated polymers such as polychlorotrifluoroethylene may also be used. For resistance to thermal and chemical degradation, the support preferably is made of a highly fluorinated polymer, most preferably perfluorinated polymer.

For example, the polymer for the porous support  
 25 can be a microporous film of polytetrafluoroethylene (PTFE) or a copolymer of tetrafluoroethylene with



30  $(m=0 \text{ to } 15, n=1 \text{ to } 15).$

Microporous PTFE films and sheeting are known which are suitable for use as a support layer. For example, U.S. Patent 3,664,915 discloses uniaxially stretched film having at least 40% voids. U.S. Patents 3,953,566,  
 35 3,962,153 and 4,187,390 disclose porous PTFE films having at least 70% voids.

Alternately, the porous support may be a fabric made from fibers of the polymers discussed above woven using various weaves such as the plain weave, basket weave, leno weave, or others.

5           A membrane can be made using the porous support by coating cation exchange polymer on the support so that the coating is on the outside surfaces as well as being distributed through the internal pores of the support. This may be accomplished by impregnating the  
10 porous support solution/dispersion with the cation exchange polymer or cation exchange polymer precursor using a solvent which is not harmful to the polymer of the support under the impregnation conditions and which can form a thin, even coating of the cation exchange  
15 polymer on the support. For example, for applying a coating of perfluorinated sulfonic acid polymer to a microporous PTFE support, a 1-10 weight percent solution/dispersion of the polymer in water mixed with sufficient amount of a polar organic solvent can be  
20 used. The support with the solution/dispersion is dried to form the membrane. If desired, thin films of the ion exchange polymer can be laminated to one or both sides of the impregnated porous support to prevent bulk flow through the membrane which can occur is large  
25 pores remain in the membrane after impregnation.

While membranes in accordance with the invention may contain a porous support and may contain inorganic filler as will be discussed in more detail hereinafter, it is preferred for the cation exchange  
30 polymer to be present as a continuous phase within the membrane.

#### Inorganic Fillers

A wide variety of inorganic fillers may be dispersed in the membranes in accordance with the  
35 invention. The inorganic filler may be essentially nonconductive such as type A zeolites of cesium, potassium, or mixtures thereof. Zeolites are a well-defined class of naturally occurring alumino

silicate minerals. They are described, e.g., by Alan Dyer in "An Introduction to Zeolite Molecular Sieves," John Wiley & Sons, Chichester, New York, 1988.

Preferably, the inorganic filler is an  
5 inorganic proton conductor which has the capability for increasing the conductivity of the membrane. Inorganic proton conductors are a known class of inorganic compounds which have the ability to transport protons. Certain classes of inorganic proton conductors are  
10 discussed, for example, by Clearfield in Solid State Ionics 46, 35 (1991). Preferred inorganic proton conductors are particle hydrates and framework hydrates. Framework hydrates are especially preferred for use in accordance with the present invention.  
15 Preferably, the metal ion of the inorganic proton conductor is selected from the group consisting of Sn, Sb, Mo, W, Ti and Zr, most preferably oxides and phosphates of these metals. Especially preferred inorganic proton conductors are titanium dioxide, tin  
20 and hydrogen mordenite, oxides and phosphates of zirconium and mixtures thereof. Preferably, the inorganic proton conductor has a conductivity of at least about  $10^{-4}$  S/cm. For the purposes of this application, the conductivity of inorganic proton  
25 conductors is measured at room temperature at 50% relative humidity.

The inorganic fillers used in the practice of this are typically commercially available, can be synthesized by known techniques, or are prepared in  
30 situ as will be described in more detail hereinafter. Hydrogen mordenite is available, for example, as an  $H^+$  mordenite powder from the PQ Corporation, Valley Forge, Pennsylvania. Tin mordenites can be prepared by ion exchange from hydrogen mordenite using the procedure  
35 described by H. Knudsen et al. in Solid State Ionics 35, 51-55 (1989).

The solubility of inorganic fillers, particularly inorganic proton conductors, varies

widely. For use in accordance with the present invention, an inorganic filler having suitably low solubility should preferably be selected to match the intended application so that the inorganic filler does  
5 not leach out of the polymer in use at a rate which would detrimentally affect performance.

The amount of inorganic proton conductor used in ion exchange membranes of this invention is preferably in the range about 2 to about 30% by weight,  
10 more preferably about 5 to about 25 % by weight, based on the total weight of the membrane.

#### Incorporation of Inorganic Fillers into the Membrane

The production of membranes containing an inorganic filler can be done by a variety of  
15 techniques. The inorganic filler can be mixed with a solution/dispersion in a suitable solvent of polymer precursor or with polymer in ionic form and cast as a film. Such solutions can also be used to apply a coatings to porous supports to form the membranes.  
20 Thermoplastic polymers, or polymers which can be in thermoplastic form, e.g., the sulfonyl fluoride form of perfluorinated sulfonic acid polymers, can be melt blended with the inorganic filler and a film can be extruded from the molten mixture. For perfluorinated  
25 sulfonic acid polymers, hydrolysis of the film to convert it to ionic form can be performed as has been discussed above although care may be required with some inorganic fillers to prevent removal or chemical alteration of the material during hydrolysis.

30 When the inorganic filler has sufficiently low water solubility, a preferred method of making membranes in accordance with the invention is to precipitate the inorganic filler *in situ* in the polymer of the membrane. This method can be used either for  
35 membranes made by film extrusion, solution film casting, or those made by coating porous supports. Preferably, for polymers such as the perfluorinated sulfonic acid polymer which are typically formed into a

film in thermoplastic form, i.e., sulfonyl fluoride form, the membrane is hydrolyzed to its ionic (sulfonate) form before *in situ* precipitation due to the greater capability of the ionic form to absorb  
5 water. For fuel cell applications, the membrane will normally be converted from an alkali metal salt form to the acid (hydrogen ion) form which may be used for the *in situ* precipitation.

*In situ* precipitation is preferably  
10 accomplished by sequentially contacting the membrane with one or more solutions containing ions or other reactants which form the inorganic filler. Using this procedure, the inorganic filler precipitates in the polymer of the membrane. In one preferred form of the  
15 process, zirconium hydrogen phosphate  $Zr(HPO_4)_2$  can be precipitated in a membrane of perfluorinated sulfonic acid polymer (preferably in acid form) by soaking the membrane in an aqueous solution of containing zirconium ions, e.g., 1-5M zirconyl chloride, for a time and at a  
20 temperature sufficient to penetrate the membrane solution. Then, the membrane is soaked in an aqueous solution containing  $(PO_4)^{3-}$  ions, e.g., 20 to 90 weight percent phosphoric acid, for a time and at a  
temperature sufficient to form zirconium hydrogen  
25 phosphate within the membrane. No special conditions are needed for carrying out the process and 2-20 hours at room temperature are suitable times for each of the soaking steps. It may be desirable to rinse the membrane in water after soaking in the zirconium  
30 solution to prevent a precipitation of filler on the surface of the membrane.

$TiO_2$  is also preferably dispersed within the polymer of a membrane by *in situ* precipitation using a process in accordance with the invention. Starting  
35 with a membrane, preferably in acid (hydrogen ion) form, a process in accordance with the invention includes soaking the membrane in an alcohol solution of one or more titanium alkoxides of the formula  $(RO)_4Ti$ .

R in this formula is an alkyl group of one to four carbon atoms. The alkoxy groups may be linear groups, such as primary alkoxy groups (e.g., propoxy), or secondary alkoxy groups such as isopropoxy and the four groups per molecule may be the same or may be different. The alcohol solvent in the titanium alkoxide solution is an aliphatic alcohol of the formula  $R'OH$  wherein  $R'$  is an alkyl group of 1 to 4 carbon atoms, for example, methanol, ethanol, propanol and butanol. The soaking is continued for a time and at a temperature sufficient to swell and expand the membrane. Preferably, temperatures in the range 20-100°C, for a sufficient time such as 1-30 minutes. A temperature of about 75°C for about 10 minutes has been found to be effective. After soaking, the surface of the membrane is preferably rinsed to wash off surface titanium alkoxide. Alcohols, such as those used for the making the titanium alkoxide solution the soaking step, are useful for rinsing. The next step of the process, is to hydrolyze the titanium alkoxide in the membrane with water. The step of hydrolyzing the titanium alkoxide in place in the membrane is carried out easily. No special conditions are required due to the ease of hydrolysis. Contact times of 10 minutes at room temperature have been found to be suitable.

In one form of the membranes in accordance with the invention and preferred membrane and electrode assemblies and fuel cells in accordance with the invention, the polymer in the bulk of said membrane is substantially free of metal catalysts. Metal catalysts such as platinum, gold, palladium, etc. have been incorporated into known membranes together with metal oxides such as  $SiO_2$  and  $TiO_2$  for the purposes of internal humidification of the membrane when used in hydrogen-oxygen fuel cells. Such catalysts are expensive and may cause undesirable effects and thus are not desirable for use in accordance with the present invention. As is explained hereinafter, metal

catalysts are typically present in membrane and electrode assemblies in which electrodes are formed on the surface of the membrane. By the bulk of the polymer being substantially free of metal catalysts is meant that the polymer in the interior of the membrane is substantially free of metal catalysts. There is no intent by this language, however, to exclude catalysts being present at or on the surface of the membrane.

10 Membrane and Electrode Assemblies and Electrochemical Cell

With reference to Figure 1, a membrane in accordance with the invention can be advantageously employed in a membrane electrode assembly (MEA) 10 in which electrodes are formed on the surface of membrane 26. MEA 10 is shown and described for use in a fuel cell although it may be used in other types of electrochemical cells which employ cation exchange membrane, especially those which transport protons across the membrane.

20 The fuel cell utilizes a fuel source indicated by arrow 12 such as methanol (typically a methanol/water solution) supplied to an anode compartment (not shown) and an oxidizer source indicated by arrow 14 such as air or oxygen supplied to an a cathode compartment (not shown). Membrane 26 serves as an electrolyte (for proton exchange) and separates the anode compartment from the cathode compartment. A porous anode current conductor 16 and a porous cathode current collector 18 are provided to conduct current from the cell. A catalyst layer 22 which functions as the cathode is in contact with and between the cathode-facing surface of membrane 26 and the cathode current collector 16. A catalyst layer 30 which functions as the anode is disposed between and is in contact with the anode-facing surface of the membrane 26 and anode current conductor 16. Cathode current collector 18 is electrically connected to



positive terminal 34 and anode current collector 16 is electrically connected to negative terminal 32.

The catalyst layers 22 and 30 may be made from well-known electrically conductive, catalytically active particles or materials and may be made by methods well known in the art. The catalyst layer 22 may be formed as a film of a polymer which serves as a binder for the catalyst particles. The binder polymer can be a hydrophobic polymer, a hydrophilic polymer or a mixture of such polymers. Preferably, the binder polymer is a polymer having cation exchange groups and most preferably is the same polymer as in the membrane. For example, in an MEA using a perfluorinated sulfonic acid polymer membrane and a platinum catalyst, the binder polymer can also be perfluorinated sulfonic acid polymer and the catalyst can be a platinum catalyst supported on carbon particles. In the catalyst layers 22 and 30, the particles are preferably uniformly dispersed in the polymer to assure that a uniform and controlled depth of the catalyst is maintained, preferably at a high volume density with the particles being in contact with adjacent particles to form a low resistance conductive path through catalyst layer.

The catalyst layers 22 and 30 formed on the membrane should be porous so that they are readily permeable to the gases/liquids which are consumed and produced in cell. The average pore diameter is preferably in the range of 0.01 to 50  $\mu\text{m}$ , most preferably 0.1 to 30  $\mu\text{m}$ . The porosity is generally in a range of 10 to 99%, preferably 10 to 60%.

The catalyst layers are preferably formed using an "ink", i.e., a solution of the binder polymer and the catalyst particles, which is used to apply a coating to the membrane. The viscosity of the ink is preferably controlled in a range of 1 to  $10^2$  poises especially about  $10^2$  poises before printing. The viscosity can be controlled by (i) selecting particle sizes, (ii) composition of the catalytically active

particles and binder, (iii) adjusting the water content (if present), or (iv) preferably by incorporating a viscosity regulating agent such as carboxymethyl cellulose, methyl cellulose, hydroxyethyl cellulose, and cellulose and polyethyleneglycol, polyvinyl alcohol, polyvinyl pyrrolidone, sodium polyacrylate and polymethyl vinyl ether.

The area of the membrane to be coated with the ink may be the entire area or only a select portion of the surface of the membrane. The catalyst ink may be deposited upon the surface of the membrane by any suitable technique including spreading it with a knife or blade, brushing, pouring, metering bars, spraying and the like. If desired, the coatings are built up to the thickness desired by repetitive application. Areas upon the surface of the membrane which require no catalyst materials, can be masked, or other means can be taken to prevent the deposition of the catalyst material upon such areas. The desired loading of catalyst upon the membrane can be predetermined, and the specific amount of catalyst material can be deposited upon the surface of the membrane so that no excess catalyst is applied. The catalyst particles are preferably deposited upon the surface of a membrane in a range from about 0.2 mg/cm<sup>2</sup> to about 20 mg/cm<sup>2</sup>.

A particularly advantageous method of applying the catalyst layers to the membrane is to use a screen printing process. It is preferable to use a screen having a mesh number of 10 to 2400, especially mesh number of 50 to 1000 and a thickness in the range of 1 to 500 mp. It is preferable to select the mesh and the thickness of the screen and control the viscosity of the ink so as to give the thickness of the electrode ranging from 1 micron to 50 microns, especially 5 microns to 15 microns. The screen printing can be repeated as needed to apply the desired thickness. Two to four passes, usually three passes, has been observed to produce the optimum performance. After each

application of the ink, the solvent is preferably removed by warming the electrode layer to about 50°C to 140°C, preferably about 75°C.

A screen mask is used for forming an electrode layer having a desired size and configuration on the surface of the ion exchange membrane. The configuration is preferably a printed pattern matching the configuration of the electrode. The substances for the screen and the screen mask can be any materials having satisfactory strength such as stainless steel, poly(ethylene terephthalate) and nylon for the screen and epoxy resins for the screen mask.

After forming the catalyst coating, it is preferable to fix the ink on the surface of the membrane so that a strongly bonded structure of the electrode layer and the cation exchange membrane can be obtained. The ink may be fixed upon the surface of the membrane by any one or a combination of pressure, heat, adhesive, binder, solvent, electrostatic, and the like. The preferred embodiment for fixing the ink upon the surface of the membrane employs pressure, heat or by a combination of pressure and heat. The electrode layer is preferably pressed onto the surface of the membrane at 100°C to 300°C, most preferably 150°C to 280°C, under a pressure of 510 to 51,000 kPa (5 to 500 ATM), most preferably 1,015 to 10,500 kPa (10 to 100 ATM).

An alternative to printing the catalyst layer directly onto the membrane is the so-called "decal" process. In this process, the catalyst ink is coated, painted, sprayed or screen printed onto a substrate and the solvent is removed. The resulting "decal" is then subsequently transferred from the substrate to the membrane surface and bonded, typically by the application of heat and pressure.

In accordance with another form of the invention, membrane and electrode assembly is provided in which its electrode comprises a layer of electrically conductive, catalytically active particles

and binder polymer having cation exchange groups with the binder polymer containing inorganic filler dispersed therein. The fillers which have been discussed previously can be suitable for incorporation  
5 into the binder polymer. Preferably, the binder polymer is the same as polymer of the membrane. Suitable polymers have been discussed previously in this patent application. In this form of the invention, it is also advantageous for the polymer of  
10 the membrane to have inorganic filler dispersed therein.

The inorganic filler can be incorporated into the binder polymer of the electrode in one of a variety of methods. One method is to incorporate the filler  
15 into the electrode ink formulation which may be directly applied to a membrane to make an MEA or may be used to form a "decal" which is subsequently applied to the membrane. A preferred method for dispersing suitable inorganic fillers in the binder polymer of an  
20 MEA is by *in situ* precipitation. Using the same process steps as are used to disperse fillers into a membrane as described previously, an inorganic filler can be dispersed in the binder polymer of the electrode of an MEA and, at the same time, dispersed in the  
25 polymer of the membrane. Similarly, an *in situ* precipitation process can be used to disperse inorganic filler in a binder polymer of an electrode "decal". Regardless of the method used to incorporate the inorganic filler in the electrode layer, care should be  
30 taken in adjusting the loading and distribution of the filler so that porosity of the electrode layer is maintained at the desired level.

The anode current collector 16 and the cathode current collector 18 may be constructed as is known in  
35 the art. These structures may be the same or different. Access of oxygen, typically air to the catalyst layer is provided by employing a porous cathode current collector 18. Similarly, the anode

current collector 16 is porous to permit the access of the methanol/water solution. While conductive metal screens, porous plates or other materials may also be used, a preferred material for the current collectors is conductive paper or cloth made of carbon fibers with suitable conductivity and porosity. Typically, the current collectors are bonded to the MEA by the application of heat and pressure.

In methanol weight gain tests, membranes in accordance with the invention with inorganic fillers dispersed therein absorb less methanol than control membranes. Similarly, a reduction in methanol transport accompanying incorporation of inorganic proton conductor is observed. It is believed that the inorganic fillers operate by filling the holes or interstices in the membrane with an organophobic material which serves to slow or block organic fuels such as methanol from diffusing through the membrane. In addition, in comparison to known membranes, preferred membranes containing inorganic proton conductors in accordance with the invention have increased proton conductivity.

The membranes in accordance with the invention are advantageously employed in MEA's for electrochemical cells, particularly fuel cells employing direct feed organic fuels such as methanol where less fuel crossover is desired. The membranes also have enhanced mechanical properties such as stiffness.

The invention is illustrated in the following examples.

#### EXAMPLE 1

This example illustrates the inorganic proton conductor,  $\text{Zr}(\text{HPO}_4)_2$ , being precipitated in situ in the polymer of a cation exchange membrane sold under the trademark NAFION® by E.I. du Pont de Nemours and Company. The membrane is 5 mil (~125  $\mu\text{m}$ ) thick and is made from a sulfonated perfluorocarbon copolymer having

a perfluorocarbon backbone and side chains of  
-O-CF<sub>2</sub>CF(CF<sub>3</sub>)-O-CF<sub>2</sub>CF<sub>2</sub>SO<sub>3</sub>H in hydrogen ion form and  
which has an equivalent weight of about 1080. The  
membrane is soaked in aqueous zirconyl chloride  
5 solution (Aldrich) at approximately 80°C for about 6  
hr. The membrane is rinsed twice with deionized water,  
and is then soaked in phosphoric acid diluted with  
water in 1:1 ratio (about 43 weight % when diluted) for  
about 12 hours. The treated membrane is finally rinsed  
10 and boiled with deionized water to remove unreacted  
phosphoric acid.

X-ray fluorescence analysis for Zr presence  
shows 1321 counts/sec, whereas the starting membrane  
shows a background of 1.9 counts/sec. Inspection of  
15 the membranes revealed that the membrane containing  
Zr(HPO<sub>4</sub>)<sub>2</sub> is significantly stiffer than the starting  
membrane.

Membranes dried at room temperature are  
additionally dried at 110°C for approximately 10 min.  
20 Liquid uptake at room temperature (about 23°C) is  
determined by measuring the weight gained after the  
dried membranes were soaked for 5 hr in deionized water  
and in dry methanol stored over molecular sieves.  
Weight gains are 20.5% for water and 40.0% for  
25 methanol. For the starting membrane, the  
corresponding weight gains are 22.0% and 54.0%,  
respectively.

Ionic (proton) conductivity is measured at 100%  
relative humidity on membranes conditioned in boiling  
30 water for about 1 hr using a Model SI 1260  
Impedance/Gain-Phase Analyzer, Solartron Instruments,  
Houston, Texas. Conductivity of the membrane  
containing Zr(HPO<sub>4</sub>)<sub>2</sub> is 0.64 S/cm, while the  
conductivity of the starting membrane is 0.40 S/cm.

35 Permeability to a 10% solution by volume of  
methanol in deionized water is measured at 65°C by  
using a Thwing Albert Cup manufactured by the Thwing  
Albert Co. For the membrane containing Zr(HPO<sub>4</sub>)<sub>2</sub> the

permeation rate is 0.0047 g/min·cm and the methanol concentration in the permeated mixture is 5.5 volume %. In contrast, for the starting membrane, permeation rate is 0.0082 g/min·cm and the methanol concentration in the permeated mixture is 8.3 volume %.

#### EXAMPLE 2

This example illustrates the inorganic proton conductor, tin mordenite, being dispersed within a membrane by using solution film casting.

A solution is prepared using water and alcohol as a solvent containing 65 wt. % of sulfonated perfluorocarbon copolymer having as a perfluorocarbon backbone and side chains of  $-O-CF_2CF(CF_3)-O-CF_2CF_2SO_3H$  in hydrogen ion form and which has an equivalent weight of about 1080. Tin mordenite is prepared from H-mordenite (supplied by PQ Corp., Valley Forge, PA under the trademark VALFOR CP 511-13) by cation exchange. The tin mordenite is added to the solution in an amount to provide 5 wt. % of tin mordenite in the resulting membrane and the solution is poured onto a polytetrafluoroethylene sheet substrate and film casting is done using a doctor blade arrangement. The resulting membrane is generally light brown in color and has a thickness of about 10-12 mil (250-300  $\mu m$ ).

Methanol and water weight gain are performed on the resulting membranes by the same method described in as used in Example 1. Water and methanol weight gain are reported in the following Table 1 compared to a control membrane made by the same solution film casting procedure but containing no tin mordenite.

Table 1

	Water(Wt. %)	Methanol(Wt. %)
1. Control	24.0	146.0
2. 5% Sn-Mordenite	22.5	117.0

Conductivity measurements are also done as detailed in EXAMPLE 1. Conductivity for the membrane containing Sn-mordenite is essentially the same as the control membrane.

5 **EXAMPLE 3**

This example illustrates the inorganic proton conductor, titanium dioxide, being dispersed within a membrane by precipitation *in situ* in the polymer of the same cation exchange membrane as used in Example 1.

10 A piece of the sulfonated perfluorocarbon polymer was cut to 54 by 74 mm (corners cut), weighing 1.105 g, was soaked in 20 ml of 1-propanol and 5 ml of  $\text{HC}(\text{OCH}_3)_3$  to remove water. It was then warmed for 10 minutes at about 50°C. Then the solution was poured  
15 off.

The polymer piece was then soaked in a mixture of 5 ml n-propanol and 1.5 ml of  $(\text{n-C}_4\text{H}_9\text{-O})_4\text{Ti}$ . After 1 hour at room temperature, an additional 2.5 ml of  $(\text{n-C}_4\text{H}_9\text{-O})_4\text{Ti}$  was added and the piece was soaked for an  
20 additional hour.

The solution was then poured off, and the piece was rinsed 1-2 seconds in room-temperature methanol. The piece was then placed in distilled water at 75°C. The piece turned stiff and milky. After 10 minutes at  
25 75-80°C, it was rinsed in cold water and wiped. It was hung and dried in air overnight. The final weight was 1.44 g, an 18% weight increase, indicating significant  $\text{TiO}_2$  loading; and hence reduced methanol migration in a fuel cell.

30



**WHAT IS CLAIMED IS:**

1. A cation exchange membrane comprising a polymer having cation exchange groups, said polymer having inorganic filler dispersed therein and wherein  
5 said polymer in the bulk of said membrane is substantially free of metal catalysts.
2. The membrane of claim 1 wherein said inorganic filler is inorganic proton conductor.
3. The membrane of claim 2 wherein said  
10 inorganic proton conductor is selected from the group consisting of particle hydrates and framework hydrates.
4. The membrane of claim 2 wherein said inorganic proton conductor is selected from the group consisting of titanium dioxide, tin and hydrogen  
15 mordenite, oxides and phosphates of zirconium, and mixtures thereof.
5. The membrane of claim 2 wherein said inorganic proton conductor has a conductivity of at least about  $10^{-4}$  S/cm.
- 20 6. The membrane of claim 1 wherein said inorganic filler comprises about 2% to about 30% by weight of said membrane.
7. The membrane of claim 1 wherein said inorganic filler comprises about 5% to about 25% by  
25 weight of said membrane.
8. The membrane of claim 1 wherein said cation exchange groups of said polymer are selected from the group consisting of sulfonate, carboxylate, phosphonate, imide, sulfonimide and sulfonamide.
- 30 9. The membrane of claim 1 wherein said polymer is highly fluorinated polymer with sulfonate cation exchange groups.
10. The membrane of claim 9 wherein said polymer comprises a highly fluorinated carbon backbone  
35 with a side chain represented by the formula  
- (OCF<sub>2</sub>CFR<sub>f</sub>)<sub>a</sub>-OCF<sub>2</sub>CFR'<sub>f</sub>SO<sub>3</sub>X, wherein R<sub>f</sub> and R'<sub>f</sub> are independently selected from F, Cl or a perfluorinated

alkyl group having 1 to 10 carbon atoms,  $a = 0, 1$  or  $2$ , and  $X$  is  $H$ , an alkali metal or  $NH_4$ .

11. A cation exchange membrane comprising a polymer having cation exchange groups, said polymer  
5 having inorganic proton conductor having a conductivity of at least about  $10^{-4}$  S/cm dispersed therein.

12. A membrane and electrode assembly comprising a cation exchange membrane of polymer having cation exchange groups with an electrode formed on at  
10 least one of its surfaces, said polymer of said membrane having inorganic proton conductor having a conductivity of at least about  $10^{-4}$  S/cm dispersed therein.

13. The membrane and electrode assembly of  
15 claim 11 wherein said electrode comprises a layer of electrically conductive, catalytically active particles.

14. The membrane and electrode assembly of Claim 11 wherein said polymer in the bulk of said  
20 membrane is substantially free of metal catalysts.

15. A membrane and electrode assembly comprising a cation exchange membrane of polymer having cation exchange groups with an electrode formed on at least one of its surfaces, said electrode comprising a  
25 layer of electrically conductive, catalytically active particles and binder polymer having cation exchange groups, said binder polymer having inorganic filler dispersed therein.

16. The membrane and electrode assembly of  
30 claim 15 wherein said polymer of said membrane is the same as said binder polymer.

17. The membrane and electrode assembly of claim 15 wherein said polymer of said membrane has inorganic filler dispersed therein.

35 18. An electrochemical cell comprising an anode compartment, a cathode compartment and a cation exchange membrane serving as a separator and electrolyte between said anode and cathode

compartments, said cation exchange membrane comprising polymer having cation exchange groups, said polymer having inorganic proton conductor having a conductivity of at least about  $10^{-4}$  S/cm dispersed therein.

5           19. A fuel cell for operation with a direct feed organic fuel, said fuel cell comprising an anode compartment containing an anode, a cathode compartment containing a cathode and a membrane serving as a separator and electrolyte between said anode and  
10 cathode compartments, and said membrane comprising a polymer having cation exchange groups and having inorganic filler dispersed therein.

          20. The fuel cell of claim 19 wherein said inorganic filler is inorganic proton conductor.

15           21. The fuel cell of claim 20 wherein said inorganic proton conductor is selected from the group consisting of particle hydrates and framework hydrates.

          22. The fuel cell of claim 20 wherein said inorganic proton conductor is selected from the group  
20 consisting of titanium dioxide, tin and hydrogen mordenite, oxides and phosphates of zirconium, and mixtures thereof.

          23. The fuel cell of claim 20 wherein said inorganic proton conductor has a conductivity of at  
25 least about  $10^{-4}$  S/cm.

          24. The fuel cell of claim 20 wherein said cation exchange groups of said polymer of said membrane are selected from the group consisting of sulfonate, carboxylate, phosphonate, imide, sulfonimide and  
30 sulfonamide.

          25. The fuel cell of claim 20 wherein said polymer of said membrane comprises highly fluorinated polymer with sulfonate cation exchange groups.

          26. The fuel cell of claim 25 wherein said  
35 polymer of said membrane comprises a highly fluorinated carbon backbone and said side chain is represented by the formula  $-(OCF_2CFR_f)_a-OCF_2CFR'_fSO_3X$ , wherein  $R_f$  and  $R'_f$  are independently selected from F, Cl or a

perfluorinated alkyl group having 1 to 10 carbon atoms,  
a = 0, 1 or 2, and X is H, an alkali metal or NH<sub>4</sub>.

27. The fuel cell of claim 19 wherein at least  
one of said anode or said cathode are formed from a  
5 layer of electrically-conductive, catalytically active  
particles on the surface of said membrane.

28. The fuel cell of claim 19 wherein the  
polymer in the bulk of said membrane is substantially  
free of metal catalysts.

10 29. The fuel cell of claim 19 wherein said  
organic fuel is methanol.

30. A process for making cation exchange  
membranes comprising dispersed inorganic filler, said  
process comprising

15 forming a membrane of polymer having cation  
exchange groups; and

dispersing inorganic filler in said membrane by  
in situ precipitation.

31. The process of claim 30 wherein said  
20 inorganic filler is titanium dioxide and said  
dispersing in said membrane is performed by:

soaking said membrane in a solution of one or  
more titanium alkoxides of the formula (RO)<sub>4</sub>Ti, wherein  
R is an alkyl group of one to four carbon atoms, said  
25 titanium alkoxides being dissolved in an aliphatic  
alcohol of the formula R'OH, wherein R' is an alkyl  
group of 1 to 4 carbon atoms, for a time and at a  
temperature sufficient to swell and expand the  
membrane; and

30 hydrolyzing with water the titanium alkoxide in  
the membrane to titanium dioxide.

32. The process of claim 30 wherein said  
inorganic filler is zirconium hydrogen phosphate and  
said dispersing in said membrane is performed by:

35 soaking said membrane in an aqueous solution of  
containing zirconium ions for a time and at a  
temperature sufficient to penetrate the membrane;

soaking said membrane in an aqueous solution containing  $(\text{PO}_4)^{3-}$  ions for a time and at a temperature sufficient to form zirconium hydrogen phosphate within said membrane.

5           33. The process of claim 30 wherein said cation exchange groups of said polymer of said membrane are selected from the group consisting of sulfonate, carboxylate, phosphonate, imide, sulfonimide and sulfonamide.

10           34. The process of claim 30 wherein said polymer of said membrane comprises highly fluorinated polymer with sulfonate cation exchange groups.

15           35. The process of claim 34 wherein said polymer of said membrane comprises a highly fluorinated carbon backbone and a side chain represented by the formula  $-(\text{OCF}_2\text{CFR}_f)_a-\text{OCF}_2\text{CFR}'_f\text{SO}_3\text{X}$ , wherein  $R_f$  and  $R'_f$  are independently selected from F, Cl or a perfluorinated alkyl group having 1 to 10 carbon atoms,  $a = 0, 1$  or  $2$ , and X is H, an alkali metal or  $\text{NH}_4$ .

20           36. The process of claim 30 wherein the bulk of said membrane is substantially free of metal catalysts.

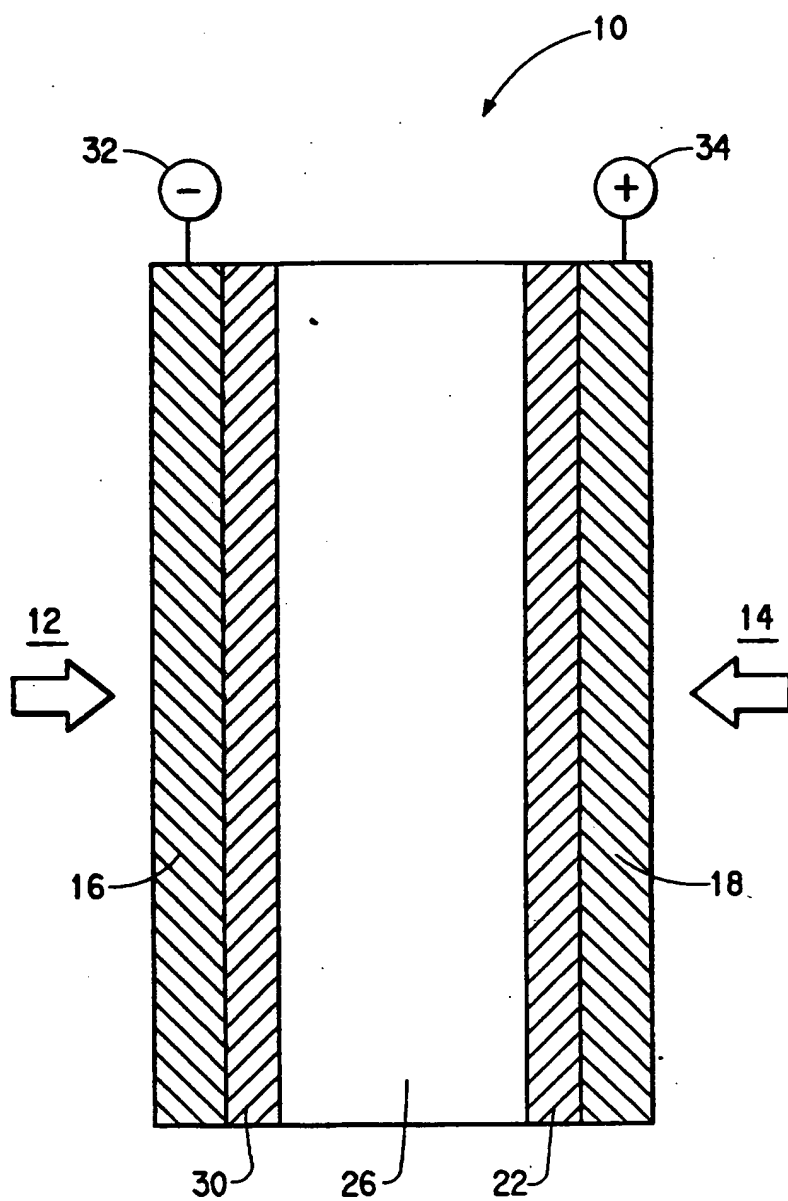


FIG. 1

## INTERNATIONAL SEARCH REPORT

Intern al Application No

PCT/US 96/03804

## A. CLASSIFICATION OF SUBJECT MATTER

IPC 6 H01M8/10 C08J5/22 H01M8/02 H01M2/16

According to International Patent Classification (IPC) or to both national classification and IPC

## B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

IPC 6 H01M C08J

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

## C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	GB,A,1 534 359 (KUREHA CHEMICAL IND CO LTD) 6 December 1978 see page 1, line 26 - page 2, line 53; claims 1-4; examples 1,7 ---	1-11,18, 30,32-36
X	US,A,4 073 752 (RAMP FLOYD LESTER) 14 February 1978 see column 5, line 52 - line 63; claims 1-3,5; examples 5,7 see abstract --- -/--	1,2, 8-10,30, 33-36

☒ Further documents are listed in the continuation of box C.☒ Patent family members are listed in annex.

## \* Special categories of cited documents:

- "A" document defining the general state of the art which is not considered to be of particular relevance
- "E" earlier document but published on or after the international filing date
- "L" document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)
- "O" document referring to an oral disclosure, use, exhibition or other means
- "P" document published prior to the international filing date but later than the priority date claimed

- "T" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention
- "X" document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone
- "Y" document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art
- "A" document member of the same patent family

Date of the actual completion of the international search

14 August 1996

Date of mailing of the international search report

02.09.96

Name and mailing address of the ISA

European Patent Office, P.B. 5818 Patentlaan 2  
NL - 2280 HV Rijswijk  
Tel. (+31-70) 340-2040, Tx. 31 651 epo nl,  
Fax (+31-70) 340-3016

Authorized officer

D'hondt, J

## INTERNATIONAL SEARCH REPORT

Intern Al Application No

PCT/US 96/03804

## C.(Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT

Category	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	CHEMICAL ABSTRACTS, vol. 90, no. 14, 2 April 1979 Columbus, Ohio, US; abstract no. 105217, MURAYAMA NAOHIRO ET AL: "Ion exchanger" XP002010753 see abstract & JP,A,53 123 390 (TOKUYAMA SODA CO) 27 October 1978 ---	1-5,8,9, 18
X	DATABASE WPI Section Ch, Week 8316 Derwent Publications Ltd., London, GB; Class A85, AN 83-38131K XP002010755 & JP,A,58 042 626 (ASAHI GLASS KK) , 12 March 1983 see abstract ---	1,6-10
X	CHEMICAL ABSTRACTS, vol. 118, no. 2, 11 January 1993 Columbus, Ohio, US; abstract no. 9408, SUZUKI NABUKAZU ET AL: "Solid-electrolyte fuel cells" XP002010754 see abstract & JP,A,04 218 267 (TOSHIBA CORP) 7 August 1992 ---	1,2,8,9, 19,20, 24-28
X	PATENT ABSTRACTS OF JAPAN vol. 011, no. 004 (E-468), 7 January 1987 & JP,A,61 181078 (HITACHI LTD), 13 August 1986, see abstract ---	1,19
X	AT,B,389 020 (SCHUETZ PETER DIPL ING DR) 10 October 1989  see page 7, line 48 - page 8, line 24; claims 1,2,12,13,17 see page 4, line 1 - line 10 ---	1,6-10, 15-17, 19,24-28
A	EP,A,0 600 470 (ASAHI GLASS CO LTD) 8 June 1994 see page 5, line 17 - line 29; claims 1,4 ---	31
X	EP,A,0 379 372 (MHB JOINT VENTURE) 25 July 1990 see page 3, line 2 - line 19; claims 1,8,11 see page 4, line 40 - line 54 see page 5, line 41 - line 51 see page 6, line 1 - line 13 ---	1,2
	---	

-/--



## INTERNATIONAL SEARCH REPORT

Intern al Application No

PCT/US 96/03804

## C.(Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	EP,A,0 206 969 (DEPARTMENT OF ENERGY MINES AND) 30 December 1986 see column 5, line 1 - line 11; claims 1,4 ---	1,2, 18-20
X	DE,A,32 10 096 (MAX PLANCK GESELLSCHAFT) 29 September 1983 see page 8, last paragraph; claims 1,6,19 ---	1,2
A	SOLID STATE IONICS, vol. 46, no. 1 / 02, 1 May 1991, pages 169-173, XP000243507 KJAER J: "SOLID STATE ELECTROLYTE MEMBRANES FOR DIRECT METHANOL FUEL CELLS" see abstract ---	4
A	US,A,3 266 940 (PHILIP D. CAESAR) 16 August 1966 see column 2, line 28 - line 45; claim 1 -----	1-4, 19-22

# INTERNATIONAL SEARCH REPORT

Information on patent family members

Insert International Application No

PCT/US 96/03804

Patent document cited in search report	Publication date	Patent family member(s)	Publication date
GB-A-1534359	06-12-78	JP-C- 1154249	30-06-83
		JP-A- 51112783	05-10-76
		JP-B- 57042090	07-09-82
		JP-C- 1292068	29-11-85
		JP-A- 52003588	12-01-77
		JP-B- 58017493	07-04-83
		JP-C- 1170483	17-10-83
		JP-A- 52008997	24-01-77
		JP-B- 58002971	19-01-83
		JP-C- 1156182	15-07-83
		JP-A- 52015487	05-02-77
		JP-B- 57050816	29-10-82
		CA-A- 1084874	02-09-80
		DE-A- 2614058	07-10-76
		FR-A,B 2306010	29-10-76
		SE-A- 7603539	01-10-76
US-A-4073752	14-02-78	AU-B- 1407576	24-11-77
		CA-A- 1048697	13-02-79
		DE-A- 2624203	23-12-76
		JP-A- 52007390	20-01-77
AT-B-389020	10-10-89	NONE	
EP-A-0600470	08-06-94	JP-A- 6263896	20-09-94
		JP-A- 6172557	21-06-94
		JP-A- 6172558	21-06-94
		US-A- 5401408	28-03-95
EP-A-0379372	25-07-90	US-A- 4990413	05-02-91
		DE-D- 69003692	11-11-93
		DE-T- 69003692	11-05-94
		ES-T- 2047838	01-03-94
		JP-A- 2236906	19-09-90
EP-A-0206969	30-12-86	CA-A- 1262041	03-10-89
		JP-A- 62003068	09-01-87
		US-A- 4724191	09-02-88
DE-A-3210096	29-09-83	JP-A- 58169868	06-10-83

**information on patient family members**

**PCT/US 96/03804**

Patent document  
cited in search report

Publication date

**Patent family member(s)**

Publication  
date

**US-A-3266940**

**16-08-66**

**NONE**